

NAME: \_\_\_\_\_ PERIOD: \_\_\_\_\_

**ACTIVITY #14 – CLASSIFYING COMPOUNDS**

**PURPOSE:**

**OBSERVATIONS:**

Compounds in Solution	Conductivity
Tap Water – H <sub>2</sub> O	
Pure Water – H <sub>2</sub> O	
Hydrochloric Acid -- HCl	
Sodium Hydroxide -- NaOH	
Sodium Chloride -- NaCl	
Barium Hydroxide – Ba(OH) <sub>2</sub>	
Sodium Nitrate – Na(NO <sub>3</sub> )	
Lead Nitrate – Pb(NO <sub>3</sub> ) <sub>2</sub>	
Sugar – C <sub>12</sub> H <sub>22</sub> O <sub>11</sub>	
Alcohol (Methyl) – CH <sub>3</sub> OH	
Potassium Chloride -- KCl	
Potassium Nitrate – K(NO <sub>3</sub> )	
Copper Chloride -- CuCl	
Copper Sulfate – Cu(SO <sub>4</sub> )	
Sodium Carbonate – Na <sub>2</sub> (CO <sub>3</sub> )	
Calcium Chloride – CaCl <sub>2</sub>	
Lead Sulfate – Pb(SO <sub>4</sub> )	

**KEY** + = Ionic, - = Covalent

## **REPORT:**

- 1) Which compounds were ionic?
- 2) Which compounds were covalent?
- 3) Which compounds conducted electricity?
- 4) Calculate the molecular or formula mass for each compound.

Compounds in Solution	Formula/Molecular Mass
H <sub>2</sub> O	
H <sub>2</sub> O	
HCl	
NaOH	
NaCl	
Ba(OH) <sub>2</sub>	
Na(NO <sub>3</sub> )	
Pb(NO <sub>3</sub> ) <sub>2</sub>	
C <sub>12</sub> H <sub>22</sub> O <sub>11</sub>	
CH <sub>3</sub> OH	
KCl	
K(NO <sub>3</sub> )	
CuCl	
Cu(SO <sub>4</sub> )	
Na <sub>2</sub> (CO <sub>3</sub> )	
CaCl <sub>2</sub>	
Pb(SO <sub>4</sub> )	

## **CONCLUSION:**